**Question 27** (2 marks)

The diagram shows a particular cell with relevant half equations.

\[
\text{Zn(s) + 2OH}^- (aq) \rightarrow \text{ZnO(s) + H}_2\text{O(l) + 2e}^- \\
\text{HgO(s) + H}_2\text{O(l) + 2e}^- \rightarrow \text{Hg(l) + 2OH}^- (aq)
\]

Identify the anode, cathode and electrolyte for this cell.

- Zn is oxidised at anode
- HgO is reduced at cathode
- Electrolyte is KOH