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~~32. a) Electrolysis of electrolysis cell. electrolysis cell.~~

~~The mixture of salt and water (Brine..~~

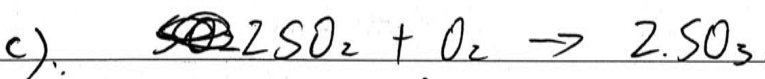
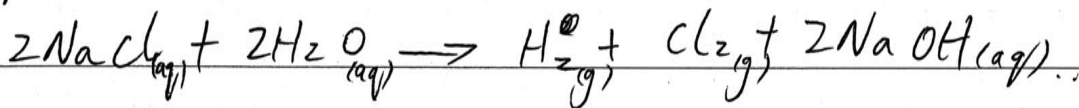
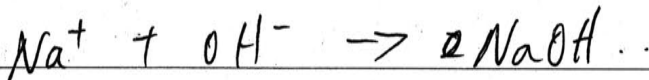
32. a). electrolysis cell.



Electrolysis the brine water that produce chlorine, hydrogen and sodium hydroxide.

process, an anode and cathode is put into the

brine water, connect the power supply, the reaction start to occur.



$$K = \frac{0.3^2}{0.5^2 + 0.25} = 0.18..$$

at A = $0.4 - \frac{0.3}{2}$

SO ₂	0.8	0.5	= 0.25 mol remain..
O ₂	0.4	0.25	
SO ₃	0	0.3	

ii) As the volume ratio is 3:2. The increases of pressure, will favour to the production side, which occur at the time B. As well, the changing temperature is another way that has a new equilibrium position. was established at time B.

32 d)

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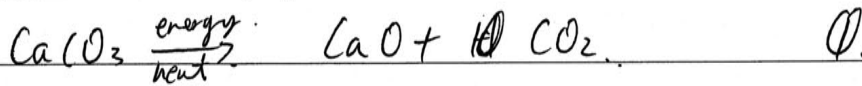
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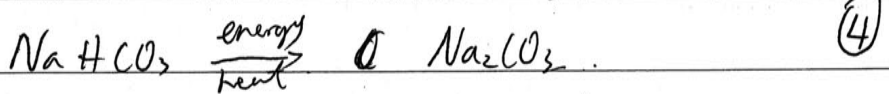
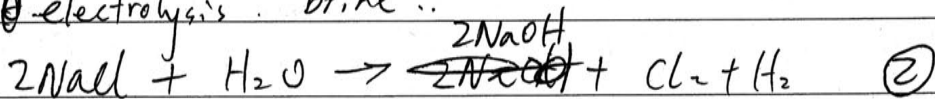
d) i) ethanal, redox.

ii) The fume is toxic, the windows should be opened. The safety glasses should be worn.

e) limestone contains $CaCO_3$.



② electrolysis brine:



Solvay process is to produce sodium carbonate. The purified carbon dioxide is needed in order to produce Na_2CO_3 through reaction ①. Limestone not only contains $CaCO_3$ but also other substances such as sulfuric. The by product of reaction ① contains sulfur dioxide, which is a pollutant and play the role in causing acidic rain; and ashes, the small ashes may be release and through into atmosphere. It endanger human's breathe system. The unsuccessful monitoring will occur smog, and become the environmental issue. As well, the production of carbon dioxide released in atmosphere that will contribute to the greenhouse effect, make globe warming even worse. It need large energy input

to ensure the reaction to occur and continues production with need. large amount of ~~electric's~~ electricity. it can also through coal burning to provide the energy for this reaction to occur, but the burning coal will produce gases such as ~~sulf~~ sulfur dioxide and Nitrogen dioxide and carbon dioxide etc.....

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