
Question 21 (3 marks)

A 0.001 mol L^{-1} solution of hydrochloric acid and a 0.056 mol L^{-1} solution of ethanoic acid both have a pH of 3.0.

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Why do both solutions have the same pH?

Both solutions have the same pH level because ~~both~~
~~ethanoic~~ the ethanoic acid is a ~~stronger~~ ^{weaker} type of acid
than the hydrochloric acid. When there is a higher ~~a~~
concentration of ethanoic acid its pH level is
raised to that of the hydrochloric acid.