

Question 26 (4 marks)

A gas is produced when 10.0 g of zinc is placed in 0.50 L of 0.20 mol L⁻¹ nitric acid. **4**

Calculate the volume of gas produced at 25°C and 100 kPa. Include a balanced chemical equation in your answer.

nitric acid: $n = cv$ $Zn + H_2NO_2 \rightarrow$
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 $n = 0.2 \times 0.5$
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 $= 0.1 \text{ mol}$
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