Question 27 (2 marks)

The diagram shows a particular cell with relevant half equations.   **2**

$$Zn(s) + 2OH^-(aq) \rightarrow ZnO(s) + H_2O(l) + 2e^-$$

$$HgO(s) + H_2O(l) + 2e^- \rightarrow Hg(l) + 2OH^-(aq)$$
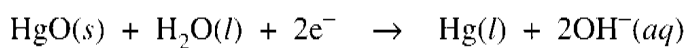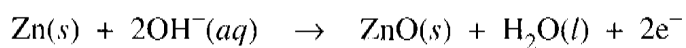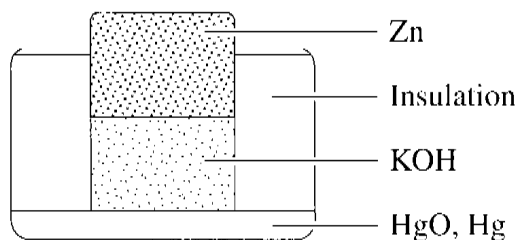
Identify the anode, cathode and electrolyte for this cell.

Zn is oxidsed at anode

HgO is reduced at cathode

electrolyte is KOH

5/6

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